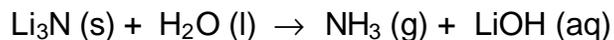


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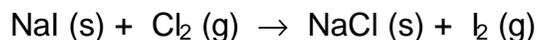
Period _____

STOICHIOMETRY WORKSHEET 1 (MASS-MASS)

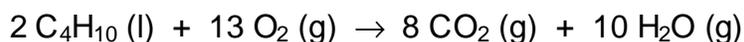
1. Determine the mass of lithium hydroxide produced when 0.38 grams of lithium nitride reacts with water according to the following **unbalanced** chemical equation:



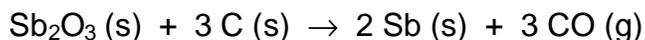
2. What mass of sodium chloride is produced when chlorine gas reacts with 0.29 grams of sodium iodide? The **unbalanced** equation is given below:



3. Determine the mass of carbon dioxide produced when 0.85 grams of butane (C_4H_{10}) reacts with oxygen according to the following balanced chemical equation:

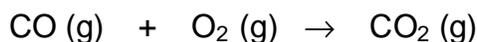


4. Determine the mass of antimony produced when 0.46 grams of antimony (III) oxide reacts with carbon according to the following balanced equation:

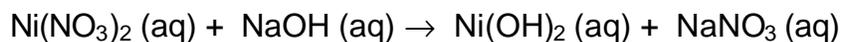


5. Hydrogen peroxide *decomposes* to produce oxygen gas and water. What mass of hydrogen peroxide (H₂O₂) must decompose to produce 0.77 grams of water?

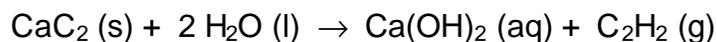
6. In the combustion of carbon monoxide, what mass of CO is required to produce 0.69 grams of carbon dioxide? The **unbalanced** equation is shown below:



7. Determine the mass of sodium nitrate produced when 0.73 grams of nickel (II) nitrate reacts with sodium hydroxide according to the following **unbalanced** chemical equation:



8. Determine the mass of calcium hydroxide produced when calcium carbide (CaC₂) reacts with 0.64 grams of water according to the following balanced chemical equation:



Answers: 1) 0.78 g LiOH 2) 0.11 g NaCl 3) 2.58 g CO₂ 4) 0.38 g Sb 5) 1.45 g H₂O₂ 6) 0.44 g CO
7) 0.68 g NaNO₃ 8) 1.32 g Ca(OH)₂