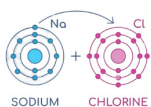


Name : \_\_\_\_\_ Date : \_\_\_\_\_



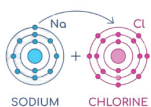
## Ionic Bonds True or False Worksheet



Identify if the following statements are true (T) or false (F).

1. An anion is positively charged. \_\_\_\_\_
2. A sodium atom needs to lose a single electron to achieve stability. \_\_\_\_\_
3. A cation usually ends with *-ide*. \_\_\_\_\_
4. Noble gases need another electron to become stable. \_\_\_\_\_
5. Halogens need another electron to become stable. \_\_\_\_\_
6. Br<sup>-</sup> is an example of a cation. \_\_\_\_\_
7. Hg<sup>2+</sup> has a configuration similar to a noble gas. \_\_\_\_\_
8. Metals become stable by transferring their electrons to become cations. \_\_\_\_\_
9. Elements from Group IIA will become stable by transferring an electron from their outermost shell. \_\_\_\_\_
10. An ionic compound always has an overall positive or negative charge. \_\_\_\_\_
11. Two metals will always form an ionic bond. \_\_\_\_\_
12. An ionic solid is always an amorphous mass. \_\_\_\_\_
13. In an electron transfer, the non-metal will transfer its electrons to a metal. \_\_\_\_\_
14. Oxygen becomes stable as an O<sup>2-</sup> ion as it has a noble gas configuration. \_\_\_\_\_
15. Helium needs eight electrons in its outermost shell to be stable. \_\_\_\_\_
16. Salts are examples of ionic compounds. \_\_\_\_\_
17. Ca<sup>2+</sup> has transferred two of its electrons to become stable. \_\_\_\_\_
18. An ionic bond is the sharing of electrons between two atoms. \_\_\_\_\_

Name : \_\_\_\_\_ Date : \_\_\_\_\_



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