

Name : Date :

Ionic Bonding of Transition Metals Worksheet

Name the following compounds.

Formula	Name of the Compound
$\text{Pb}_3(\text{PO}_4)_2$	
$\text{K}_2\text{Cr}_2\text{O}_7$	
$\text{Ba}(\text{ClO}_4)_2$	
$\text{Fe}(\text{NO}_2)_2$	
$\text{Fe}(\text{C}_2\text{H}_3\text{O}_2)_2$	
$(\text{NH}_4)_2\text{CrO}_4$	
$\text{Hg}_2(\text{NO}_2)_2$	
$\text{Pb}(\text{Cr}_2\text{O}_7)_2$	
$\text{Al}_2(\text{Cr}_2\text{O}_7)_3$	
$(\text{NH}_4)_2\text{C}_2\text{O}_4$	
$\text{Ag}_2(\text{Cr}_2\text{O}_7)$	
$\text{Pb}(\text{MnO}_4)_4$	
$\text{Pb}(\text{ClO}_3)_4$	
BaCr_2O_7	
CsMnO_4	
CuClO_3	

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Name the following compounds.

Formula	Name of the Compound
$\text{Pb}_3(\text{PO}_4)_2$	Lead(II) Phosphate
$\text{K}_2\text{Cr}_2\text{O}_7$	Potassium Dichromate
$\text{Ba}(\text{ClO}_4)_2$	Barium Perchlorate
$\text{Fe}(\text{NO}_2)_2$	Iron (II) Nitrite
$\text{Fe}(\text{C}_2\text{H}_3\text{O}_2)_2$	Iron (II) Acetate
$(\text{NH}_4)_2\text{CrO}_4$	Ammonium Chromate
$\text{Hg}_2(\text{NO}_2)_2$	Mercury (I) Nitrite
$\text{Pb}(\text{Cr}_2\text{O}_7)_2$	Lead (IV) Dichromate
$\text{Al}_2(\text{Cr}_2\text{O}_7)_3$	Aluminum Dichromate
$(\text{NH}_4)_2\text{C}_2\text{O}_4$	Ammonium Oxalate
$\text{Ag}_2(\text{Cr}_2\text{O}_7)$	Silver Dichromate
$\text{Pb}(\text{MnO}_4)_4$	Lead (IV) Permanganate
$\text{Pb}(\text{ClO}_3)_4$	Lead (IV) Chlorate
BaCr_2O_7	Barium Dichromate
CsMnO_4	Cesium Permanganate
CuClO_3	Copper (I) Chlorate